

# **Welcome to a Water-Filled Friday of Science Class!**

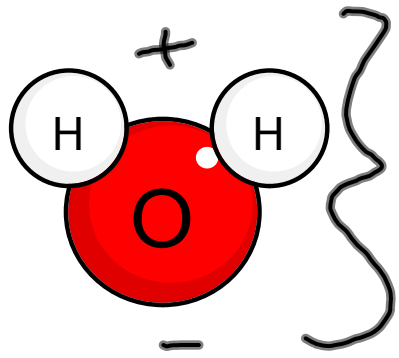
**Find your lab paper from this week, over on the lab tables at the side of the room. Then find your assigned seat from Wednesday, and get ready to do some learnin'!**

## **Today's Agenda**

- **Post-Lab Discussion**
- **Lesson: Water's Polarity & Related Properties**
- **HW:**
  - 1. Unit 3 Preassessment (on Moodle), for completion grade**
  - 2. 'Properties of Water' Lab Analysis Questions (on Moodle)**

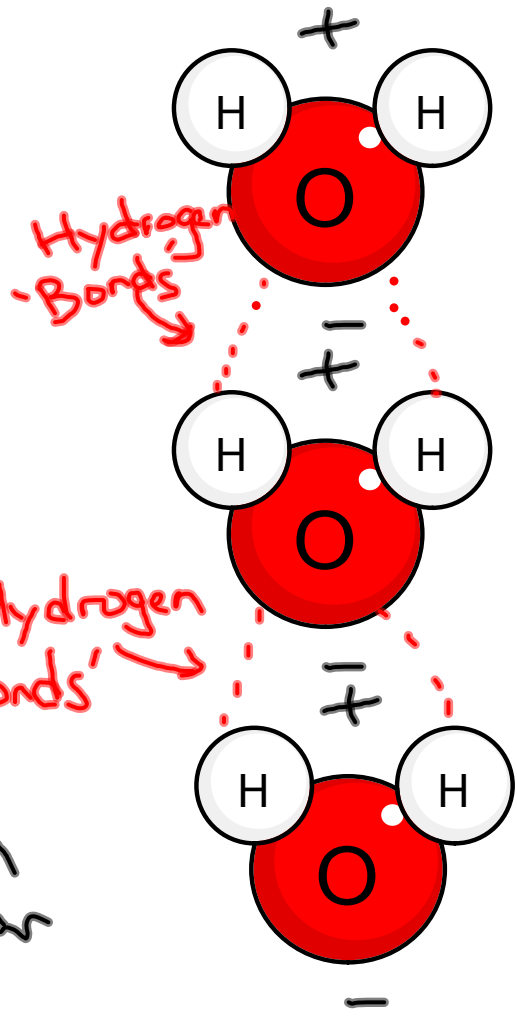
**Both Moodle HW assignments are due by Mon @ 7:30 am.**

# H<sub>2</sub>O Molecules

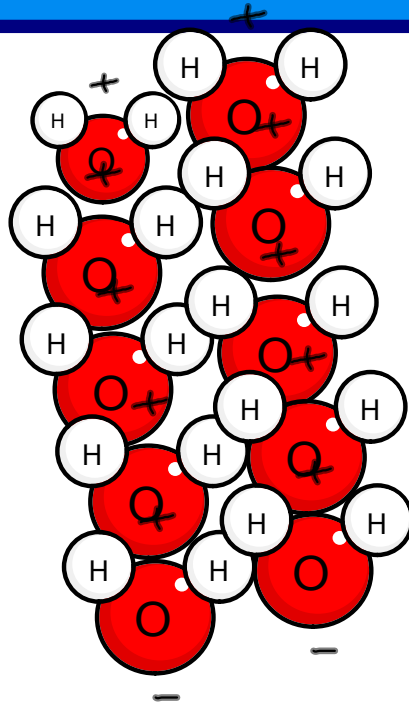
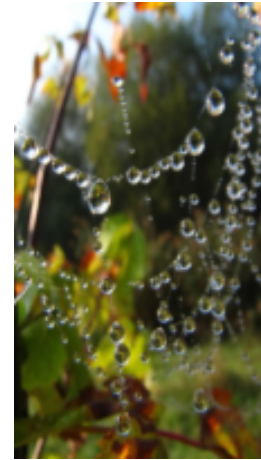


Water is polar because its ends have opposite charges.

Hydrogen 'Bonds' are relatively weak attractions between opp. charges in polar molecules



# Cohesion & Surface Tension



Cohesion - 'sticking' together of many molecules of the same type  
(Adhesion is when different substances stick together.)

Surface Tension - "sticky skin" that forms @ surface of liquid  
- property of a liquid that determines how much the surface behaves like elastic.  
\* Water has a very high surface tension.

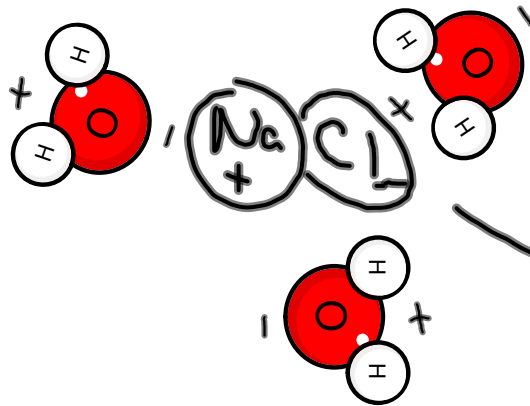
## Water as a Solvent

\* When a substance dissolves, its particles are breaking into smaller pieces and spreading apart from another.

Solvent - substance that makes the solute dissolve

Solute - substance that dissolves

\* "Like dissolves like."



Water's good @ dissolving salt, because salt is polar.